GEOCHEMISTRY

Photoferrotrophy, deposition of banded iron formations, and methane production in Archean oceans

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Banded iron formation (BIF) deposition was the likely result of oxidation of ferrous iron in seawater by either oxygentic photosynthesis or iron-dependent anoxygenic photosynthesis—photoferrotrophy. BIF deposition, however, remains enigmatic because the photosynthetic biomass produced during iron oxidation is conspicuously absent from BIFs. We have addressed this enigma through experiments with photosynthetic bacteria and modeling of biogeochemical cycling in the Archean oceans. Our experiments reveal that, in the presence of silica, photoferrotrophic cell surfaces repel iron (oxyhydroxides). In silica-rich Precambrian seawater, this repulsion would separate biomass from ferric iron and would lead to large-scale deposition of BIFs lean in organic matter. Excess biomass not deposited with BIF would have deposited in coastal sediments, formed organic-rich shales, and fueled microbial methanogenesis. As a result, the deposition of BIFs by photoferrotrophs would have contributed fluxes of methane to the atmosphere and thus helped to stabilize Earth’s climate under a dim early Sun.

INTRODUCTION

Banded iron formations (BIFs) host the world’s largest iron ore deposits, and they formed predominantly through the deposition of ferric iron [Fe(III)] from ferruginous oceans during the Archean Eon (1, 2). Most models for BIF deposition invoke photosynthesis in the oxidation of ferrous iron from seawater to induce its subsequent precipitation and deposition as mixed valence iron (oxyhydr)oxides and carbonate phases (3–6) [see (7) and (8) for alternative models]. BIFs thus likely record the activity of Earth’s early photosynthetic biosphere (3). Two modes of photosynthesis have been implicated in Fe(II) oxidation—canonical oxygentic photosynthesis by the ancestors of modern cyanobacteria (3) and iron-dependent anoxygenic photosynthesis (4, 5), referred to as photoferrotrophy. Photoferrotrophic bacteria grow using light and Fe(II) to fix CO₂ into biomass and produce Fe(III) as a metabolic by-product (5, 6, 9)—they can do so in the complete absence of oxygen (9). Current literature suggests that this anoxygenic photosynthesis is likely the evolutionary predecessor of oxygentic photosynthesis (10–12) and that Fe(II) oxidation could thus have been driven by photoferrotrophy before the emergence and proliferation of oxygenic photosynthesis (13, 14). Models of nutrient cycling in the Archean ocean implicate photoferrotrophs as key primary producers (13, 15) before the rise of atmospheric oxygen 2.4 to 2.3 billion years (Ga) ago during the Great Oxidation Event (GOE) (16–18). High methane concentrations in the Archean atmosphere have also been qualitatively attributed to C and Fe cycling associated with BIF deposition (19). Such an “upside-down” biosphere (19), in which the reduced products of photosynthesis end up in the atmosphere as methane and the oxidized products are buried as Fe(III) in BIF, can, as we show here, be quantitatively and mechanistically linked to the activity of photoferrotrophs in ferruginous Archean oceans.

Observations and models of extant photoferrotrophs demonstrate their capacity to both oxidize Fe(II) at rates sufficient to form even the largest BIFs and support appreciable rates of primary production (5, 13, 20, 21). Laboratory experiments to date, paradoxically, reveal a tight physical association between photosynthetic ferric iron metabolic by-products [e.g., Fe(III) (oxyhydr)oxides] and cellular biomass that leads to their co-sedimentation (5, 20, 22). In the oceans, the deposition of Fe(III), along with this photosynthetic biomass, would fuel sedimentary respiration that couples oxidation of organic carbon to microbial Fe(III) mineral reduction and converts Fe(III) to dissolved Fe(II) and secondary minerals such as siderite and magnetite (22). This sedimentary respiration thus closes the iron redox cycle and precludes the preservation of ferric iron in BIF. Current observations and models therefore cannot explain BIF deposition through photoferrotrophy in light of this co-sedimentation of biomass with Fe(III) and ensuing diagenetic reactions (22, 23). By contrast, a lack of Fe(III) reduction would lead to the co-deposition of ferric iron in BIF, with organic matter at relative concentrations close to the 4:1 stoichiometry of Fe(II) oxidation to C fixation in photoferrotrophy [e.g., ~2 weight % (wt %) C based on 37 wt % Fe(III) in Dales Gorge BIF (24)]. A compilation of organic matter concentrations and Fe redox states in BIFs, however, reveals that they have very low organic matter concentrations (mean of 0.27 wt %; Fig. 1 and table S5) (24, 25) and contain appreciable Fe(III), with an average Fe redox state of 2.6 (fig. S1 and table S5) (24, 25). If the iron minerals in BIFs were the product of photosynthesis, then the fate of the corresponding photosynthetic biomass remains entirely enigmatic. Furthermore, regardless of the mode of Fe(II) oxidation, current estimates for the magnitude of hydrothermal Fe(II) fluxes to the Archean ocean appear to be deficient of the mass fluxes needed to sustain BIF deposition (26) and may imply an important role for terrestrial weathering fluxes (27, 28). We thus combined experiments

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using modern photosynthetic bacteria with revised theoretical estimates of Fe(II) fluxes to the Archean ocean and box modeling of coupled carbon and iron cycling to show that photosynthetic Fe(II) oxidation could have sustained large-scale, organic-poor BIF deposition in ocean upwelling systems over hundreds of millions of years.

RESULTS AND DISCUSSION
Separation of biomass and Fe(III)
To explore the fate of biomass and Fe(III) during photosynthetic Fe(II) oxidation, we conducted experiments using *Chlorobium phaeoferrooxidans* strain KB01—a photoferrotrophic bacterium that was isolated from the water column of ferruginous Kabuno Bay, a sub-basin of Lake Kivu in East Africa (21). We also conducted experiments with *Chlorobium* ferronitrificans strain KoFox, a sediment-dwelling photoferrotroph that tends to associate with Fe(III) (oxyhydr)oxides without becoming encrusted (23, 29), as well as a model marine oxyphototrophic cyanobacterium of the genus *Synechococcus*. Strain KB01 is the only known photoferrotroph from a pelagic environment (21), and pelagic photoferrotrophs, including those that would have populated Precambrian ferruginous oceans, face a special challenge—they must maintain their position within sunlit waters despite the precipitation of heavy Fe(III) (oxyhydr)oxides, leading to co-sedimentation of biomass and Fe(III) (Fig. 2A). While phosphate concentrations were very low in the seawater from which BIFs deposited (15, 33), silica concentrations in Precambrian oceans were high (~1 mM) (15), and this would have altered the physical and chemical properties of Fe(III) precipitates formed in seawater (33). In experiments with 1 mM silica, the concentration implied for Precambrian seawater (15), strain KB01 cells did not associate to, or co-sediment with, their Fe(III) (oxyhydr)oxide by-products (Fig. 2A). Instead, almost all of the cells remained suspended (94 ± 6%). Likewise, experiments conducted with strain KoFox demonstrated that cell-mineral association was diminished relative to the silica-free experiments, with 72 ± 7% of strain KoFox cells remaining suspended (Fig. 2A and fig. S5A). We also grew both strains KB01 and KoFox in their standard growth medium containing 4 mM phosphate and no silica, which led to little cell-mineral association for strain KB01 (Fig. 2A) and modest association for strain KoFox (fig. S5A). Detailed electron microscopy revealed that the surfaces of KB01 cells were entirely free from Fe(III) (oxyhydr)oxide precipitates (Fig. 3), whereas strain KoFox tended to form multicellular aggregates that variably associated with the mineral precipitates (fig. S3A). Strain KB01, and to a somewhat lesser extent strain KoFox, thus avoids encrustation with Fe in low-P, high-Si waters, as well as in standard growth media with high phosphate, remaining suspended despite Fe(III) (oxyhydr)oxide precipitation.

We also subjected *Synechococcus* cells to Fe-Si–rich growth conditions in a similar fashion (Fig. 2A and fig. S5A). *Synechococcus* associated to a greater extent with the Fe(III) (oxyhydr)oxide products of Fe(II) oxidation than did the photoferrotrophs under these conditions. Such an association was expected, despite the presence of silica, given that oxygen effuses from *Synechococcus* cells, reacts rapidly with Fe(II), and causes precipitation of Fe(III) (oxyhydr)oxides on the cell surfaces—an association likely maintained due to binding with organic ligands, as observed in many previous environmental and laboratory studies (32, 34, 35). Photoferrotrophs adapted to pelagic lifestyles thus appear most capable of avoiding co-sedimentation with Fe(III) and remain buoyant with the potential for separation of biomass from Fe(III) at larger scales in the environment.

Mechanisms of cell-mineral separation
Surface charge often influences cell–mineral association, or lack thereof, and as solution P:Fe or Si:Fe ratios increased, the surface charge on the Fe(III) (oxyhydr)oxide formed became more negative (Fig. 2B) due to the incorporation of P or Si anions into the oxyhydroxide structure (33). Conversely, Fe(III) (oxyhydr)oxide particles were positively charged under low P or Si conditions, which is consistent with observations of positively charged Fe(III) precipitates in many environments (31). Strain KB01 cells had strongly negative surface charges (fig. S4), whereas cells of strain KoFox had near neutral surface charges (fig. S4), similar to some micro-aerophilic Fe(II) oxidizers (36). This difference in surface charge between strains KB01 and KoFox reflects different cell surface chemistries (table S2), notably an abundance of anionic surface functional groups on strain KoFox with different acid–base behavior than strain KB01 (Supplementary Materials and fig. S3). Benthic microorganisms, such as strain KoFox, commonly produce surface
Fig. 2. Cell surface characteristics for strain KB01 and the relationship between ferric iron surface charge and medium anions. Fractions of planktonic (blue) versus sedimented (red) cells (A) for photoferrotrophic strain KB01 under varying geochemical conditions: 400 µM Fe(II) with low P (3 µM) [unmarked]; 400 µM Fe(II), low P (3 µM), with 0.6 mM Si [A]; 400 µM Fe(II), low P (3 µM), with 1.0 mM Si [B]; 400 µM Fe(II), low P (3 µM), with 1.5 mM Si [C]; and 10 mM Fe(II) with 4.4 mM P [D]. *C. ferrooxidans* and *Synechococcus* are also shown under the [C] conditions. Rhodobacter strain SW2 is shown under the [D] conditions. The zeta potential (B), in millivolts, of the ferric iron precipitates is depicted with an increasing ratio of Si or P to Fe, where the error bars are all within the data points. Last, for the extended DVLO modeling for strain KB01 (C), the main graph depicts the interaction energies of the three forces (AB, Lewis acid base; LW, Lifshitz–van der Waals; EL, electrostatic) and the total (TOT) for those forces from 3 to 5 nm, while the inset depicts the forces from 0 to 5 nm.

layers rich in anionic functional groups to facilitate attachment to solid substrates (37), and these anionic surface functional groups tend to strongly bind Fe(III), effectively neutralizing surface charge (fig. S4). The role of Fe(III) in controlling cell-mineral association, and ultimately like-charged surfaces cause the physical separation of strain KB01 and, by extension, other pelagic photoferrotrophs from their Fe(III) oxyhydroxide by-products. These findings have important implications for BIF deposition and the coupling of carbon and iron cycles in the Archean Eon. We note that while strain KB01 is a derived member of the phylum *Chlorobi*, its cell surface chemistry is typical for Gram-negative bacteria (30, 39), and our results are thus likely extensible to its ancestors and most other Gram-negative bacteria, with the exception of benthic organisms, like strain KoFox, that have unusual modifications to their cell surfaces. Stem group, pelagic photoferrotrophs in Archean oceans, therefore, most likely had cell surface chemistries and interacted with Fe(III) (oxyhydr)oxides much like strain KB01.

**Revised Precambrian Fe budgets**

Before proceeding to model coupled carbon and iron cycling, we sought to reconcile the possible material fluxes of Fe(II) to the oceans and those needed to sustain BIF deposition (26, 40). These material fluxes serve as boundary conditions for our models and help tether our results to the geologic record. Peak BIF deposition rates imply an Fe burial flux of 45 mol m⁻² year⁻¹ over areas greater than 10¹¹ m², or ~4.5 Tmol Fe year⁻¹ (5), and at steady state, Fe(II) must
be actively resupplied to the oceans at this rate through a combination of hydrothermal venting and continental and seafloor weathering. Previous estimates of modern hydrothermal Fe fluxes of ~2 Tmol year\(^{-1}\) (Supplementary Materials), on their own, are insufficient to support steady-state BIF deposition (table S3) (1, 26, 40). These estimates were based on the product of hydrothermal fluid flow and Fe(II) concentrations measured in circulating on- and off-axis fluids (41). Global Fe fluxes, however, would have been much different in the Archean Eon because of enhanced hydrothermal activity (40, 42), smaller continents (43), reduced seawater sulfate (42, 44), weathering of Earth’s crust at low oxygen (18, 45, 46), and pervasively anoxic oceans.

Hydrothermal fluid fluxes to the oceans have been calculated from a variety of geochemical and physical data resulting in a wide range of estimates for these fluxes in the literature. A recent compilation from a number of sources (47), however, provides a synthesis of both on- and off-axis fluid flow as well as a modeled distribution of these values that yields averages (modes) of \(5 \times 10^{13}\) and \(1.5 \times 10^{14}\) kg year\(^{-1}\), respectively (47). These values are greater than values considered in previous estimates of modern hydrothermal Fe(II) budgets, implying that modern Fe(II) fluxes to the oceans were underestimated (41). Archean hydrothermal fluid flow was, furthermore, likely higher than today due to greater heat loss from the Archean lithosphere (26, 48). Following previous work (26), we thus scaled the updated estimates for modern fluid flow by the ratio of past to modern lithospheric heat loss (Supplementary Materials). This yields estimates for hydrothermal fluid flow at 2.5 Ga of \(1 \pm 0.6 \times 10^{14}\) and \(4 \pm 0.6 \times 10^{16}\) kg year\(^{-1}\) for on- and off-axis venting, respectively (table S3). Crucially, these revised estimates for fluid flow increase the possible Fe(II) fluxes to both the modern and Archean oceans.

The concentration of Fe(II) in modern on-axis hydrothermal fluids is ~6 mmol kg\(^{-1}\) (41), but this depends on the chemistry of seawater that circulates through seafloor basalts. Notably, high concentrations of sulfate (28 mM) in modern seawater lead to production of hydrogen sulfide in anoxic hydrothermal fluids, and this hydrogen sulfide reacts with Fe(II) to form iron-sulfide minerals that limit the concentration of Fe(II) in effluxing fluids (42). Sulfate concentrations in Archean seawater were much lower than today due, in part, to the lack of oxidative pyrite weathering on the continents (44). Models of hydrothermal fluid chemistry, therefore, predict that Fe(II) concentrations in high-temperature hydrothermal fluids were 10-fold higher when seawater was sulfate free versus the modern 28 mM (42), implying that high-temperature vent fluids in the Archean Eon could have had up to 60 mmol kg\(^{-1}\) Fe(II). Combining such a high-temperature hydrothermal Fe(II) concentration (42) with revised on-axis hydrothermal fluid flows results in Fe(II) fluxes of \(8 \pm 3\) Tmol year\(^{-1}\) at 2.5 Ga (table S3), and these are similar to, but smaller than, other recent estimates (49). The effect of seawater sulfate concentrations on the Fe(II) concentration in lower-temperature, off-axis hydrothermal fluids appears less pronounced, and so we conservatively consider off-axis hydrothermal Fe(II) concentrations equivalent to today (~0.75 mmol kg\(^{-1}\)) (42). This results in possible off-axis Fe(II) fluxes of up to \(30 \pm 5\) Tmol year\(^{-1}\) (table S3). Combining these new estimates for on- and off-axis Fe(II) fluxes, we redefine upper possible limits on hydrothermal Fe(II) delivery to the oceans at 2.5 Ga as ~40 Tmol year\(^{-1}\) (table S3).

In addition to the Fe(II) fluxes from hydrothermal venting, modern continental weathering also contributes Fe to the oceans. We thus assessed the potential delivery flux of continental Fe(II) to Archean oceans, taking into consideration the low oxygen atmosphere (45) and smaller continent sizes (43), resulting in a total Fe(II) weathering flux of 5 Tmol year\(^{-1}\) to the oceans (Supplementary Materials). Summing continental weathering and hydrothermal Fe(II) fluxes yields an upper possible global Fe(II) flux of ~45 Tmol year\(^{-1}\) to the oceans 2.5 Ga (table S3). This is more than sufficient to deposit BIF at 4.5 Tmol year\(^{-1}\) and is supplied mostly through previously unconsidered off-axis hydrothermal venting. These revised Fe(II) fluxes place upper boundaries on the magnitude of global coupled carbon and iron cycling, which we explore below using biogeochemical box models.

**Modeling Archean marine iron and carbon cycles**

We assume that biological production in our model is driven by photoferrotrophy, and while oxygenic photosynthesis could have been active from the Mesoarchean (46, 50), its contribution to primary production may have been small before the GOE due to competition between modes of photosynthesis would have no real effect on our model outcomes because oxygen produced would react with Fe(II), yielding the same 4:1 stoichiometry between iron oxidation and carbon fixation as in photoferrotrophy. Biomass degradation in our model was first channeled through heterotrophic Fe(III) reduction and then, given the near absence of sulfate in the Archean ocean (44), microbial methanogenesis (22, 44). We implicitly assumed that both Fe(III) reduction and methanogenesis are preceded by the breakdown of organic matter through hydrolysis, glycolysis, and fermentation. Rates of organic matter breakdown were parameterized on the basis of observations from modern anoxic marine basins (52, 53). However, we note that, unlike today, the biological pump in the Archean oceans would have operated without ballasting from fecal pellets. Biological production and nutrient cycling in the modern oceans can be divided into three broad oceanographic provinces—open ocean, coastal zone, and upwelling regions (52). In detail, the distribution and biological activity of these provinces depend on ocean circulation patterns, continental configurations, and nutrient supply, but in the absence of robust constraints on these parameters for the Archean oceans, we assumed that the relative contributions of similar such provinces to biological production scale with continental area and were otherwise similar to the modern. We thus distributed total biological production across these three provinces according to their relative productivities in the modern ocean (52) and scaled continental area from 0 to 100% of the modern (Fig. 4 and Supplementary Materials).

Our results revealed that upwelling provinces support Fe(II) oxidation and Fe(III) oxyhydroxide sedimentation at rates sufficient to deposit even the largest BIFs, such as those of the Hamersley Basin in Western Australia (Fig. 5, Table 1, and Supplementary Materials). Notably, Fe(III) (oxyhydr)oxides, with a relatively high density (3.8 g cm\(^{-3}\)) and a tendency to aggregate to larger particle sizes (54), had an average settling velocity of \(2 \times 10^{4}\) m year\(^{-1}\). This led to their deposition within a maximum of 40 km from their locus of initial precipitation, assuming a 150-m-deep water column and current velocities less than 650 m year\(^{-1}\) (fig. S6 and table S4) (52, 55). Furthermore, this localized deposition occurred despite the strong horizontal current velocities that are characteristic of upwelling provinces (52, 55). In stark contrast, the average settling velocity of biomass not associated with Fe(III) (oxyhydr)oxides was 0.35 m year\(^{-1}\), and biomass can thus be transported at distances more than 6000 km—greater than
the width of the modern Pacific Ocean (fig. S6 and table S4). Substantial biomass can thus be exported from productive upwelling areas and broadly distributed between oceanic provinces before its deposition. Notably, the separation of biomass from Fe(III) (oxyhydr) oxides led to little pelagic and sedimentary Fe recycling within the upwelling province itself. The small fraction of biomass deposited in upwelling regions fueled diagenesis, resulting in both Fe(II) recycling to the water column and the authigenesis of reduced Fe phases such as siderite and magnetite. Siderite can be an abundant component in many BIF (24, 25), and an example carbon isotope mass balance suggests that 30% of the siderite in BIFs can be diagenetic, with the balance likely representing a primary precipitate (Supplementary Materials). In our benchmark scenario (Supplementary Materials), where 15% of cells are associated with their Fe(III) by-products, organic carbon deposition rates could have supported the conversion of 10% of the total Fe(III) deposited to diagenetic siderite. When the diagenetic siderite is combined with primary nondiagenetic siderite, as inferred from the isotopic composition of siderite, the total siderite content of BIFs in our benchmark scenario is as high as 30%. This combination of primary and diagenetic siderite and remaining Fe(III) (oxyhydr)oxides yielded an average redox state of Fe in BIF of 2.7 and is very similar to that of many siderite-bearing Neoarchean BIFs (fig. S1 and table S5).

Biomass exported from upwelling provinces augmented the biomass produced in the other provinces and drove extensive pelagic and sedimentary Fe recycling through heterotrophic Fe(III) reduction. As decoupling of biomass from Fe(III) (oxyhydr)oxides in upwelling regions led to excess biomass over Fe(III) in both other provinces, it fueled methanogenesis once Fe(III) had been entirely reduced or was buried as Fe-poor, organic carbon–bearing coastal and deep-sea sediments (fig. 4). The deposition of biomass in these sediments would have led to organic carbon concentrations of between 0.5 and 5 wt % depending on sedimentation rates of detrital material (Supplementary Materials and fig. S7). At low sedimentation rates characteristic of deep-water environments, organic carbon concentrations are similar to those of the organic carbon–rich shales that deposited throughout the Archean Eon (45).

The globally integrated rates of these processes varied depending on continental area, biomass association to Fe(III) (oxyhydr)oxides, Fe(II) concentrations in the deep ocean, and rates of upwelling (Fig. 5). These variables can be constrained to likely ranges based on material fluxes recorded in sedimentary rocks (1, 2, 40), bounds for global cycling (1, 28, 40), and tempered analogies to modern systems (21, 56)—model outputs within these ranges are delineated by the shaded area in Fig. 5. Deep ocean Fe(II) concentrations are constrained to <70 µM, similar to earlier estimates of 50 µM based on siderite and calcite solubility (28), because higher Fe(II) concentrations led to global Fe(II) oxidation rates that exceed the revised maximum global Fe(II) supplies and recycling combined (table S3). Even at these relatively low Fe(II) concentrations (20 to 70 µM), upwelling provinces could have supported area-specific rates of BIF deposition of up to 54 mol m−2 year−1 over areas greater than 1013 m2 (Table 1), as needed to deposit the largest of the Archean BIFs (5).

Notably, with decreased cell-mineral association, increased continent size, increased deep-water Fe(II) concentrations, and increased upwelling rates, Fe(II) consumption could have eventually outpaced even upper limits for Fe(II) supply, and this would have led to depletion of deep ocean Fe(II). Within the most likely parameter space, rates of global primary production through photoferrotrophy would have been less than 1% of modern primary production (Table 1) and, in the absence of oxygen and sulfate, most of this
The model prediction that methanogenesis would have played a key role in carbon degradation in Archean ferruginous oceans is also supported by other Earth system models (49) and observations from modern ferruginous basins. In Kabuno Bay, East Africa, for example, carbon produced through photoferrotrophy is remineralized through methanogenesis (21), leading to an export of ferric iron from the euphotic zone even under extremely quiescent physical conditions. Likewise, despite abundant Fe(III) (oxyhydr)oxides, as much as 50% of the carbon degradation in ferruginous Lake Matano is channeled through methanogenesis, rather than Fe(III) reduction (56). Empirical observations thus point generally to carbon and iron cycling under ferruginous conditions that ultimately leads to methane production even in the absence of the large-scale advective processes that operate in the oceans. When integrated into a box model of the Archean marine iron and carbon cycles, our observations of cell-Fe(III) oxyhydroxide separation during photoferrotrophy quantitatively describe an Archean biosphere that: (i) accounts for the Fe(III) deposition rates needed to form BIFs; (ii) explains the fate of biomass and its absence from BIFs; and (iii) provides a strong source of methane to the biosphere.

This model demonstrates that, within the geologically constrained parameter space, global rates of methanogenesis can easily reach 3.2 Tmol year\(^{-1}\) (Table 1), and with the near absence of oxygen and sulfate to fuel methane oxidation, most of this methane would be delivered to the atmosphere (58). On the basis of new solutions (49) to photochemical models (59), this biospheric methane flux (3.2 Tmol year\(^{-1}\)) would support a 10 ppmv (parts per million by volume) methane atmosphere under our benchmark scenario, with a possible range of between 1 and 20 ppmv across the likely model parameter space (Table 1). While these methane concentrations alone are unlikely to support a warm climate, positive feedbacks with other concurrent modes of photosynthesis, like H\(_2\)-based anoxygenic photosynthesis, markedly increase both biospheric methane fluxes to the atmosphere and atmospheric methane concentrations (49). Our observations of extant photoferrotrophs, revised global Fe budgets, and models of coupled C and Fe cycling thus support an upside-down Archean biosphere, similar to that originally proposed by Walker (19), in which the deposition of BIFs leads to a methane-rich atmosphere. Our results further imply that cell-Fe(III) separation augments marine methane production through the breakdown of photoferrotrophic biomass and mechanistically ties biospheric methane fluxes to the deposition of BIF. The role this plays in atmospheric chemistry and greenhouse warming should be tested through further biogeochemical-photochemical modeling efforts.

**METHODS**

Photoferrotrophic strains and cyanobacteria were grown in basal media (21) and in alternative media (Supplementary Materials and table S1) until late exponential phase. Subsamples of both Fe(II)/Fe(III) and pigments were used to track the growth kinetics and to assess the in vitro cellular association to Fe(III). The surface properties of each strain and their metabolic Fe(III) by-products were assessed by measuring their surface potential using the Particle Metrix ZetaView. The acid-base properties of cell surfaces were determined through titrations. Cell surface contact angles were measured for DVLO modeling in three different liquids—water, glycerol, and diiodomethane (36, 38). Electron microscopies [scanning electron microscopy (SEM) and transmission electron microscopy (TEM)] were used to image cell-mineral interactions (Supplementary Materials). The particle sizes of the Fe(III) (oxyhydr)oxides produced by the photoferrotrophic strains were measured using a Mastersizer 2000 particle size analyzer, and these were then used to model the impact of horizontal ocean
current velocities on the settling time of both Fe(III) particles and cellular biomass (Supplementary Materials, fig. S6, and table S4). The box model of the Archean marine carbon and iron cycles was designed with three boxes (each representing an oceanic province) to capture variability in fluxes through each box (Supplementary Materials, Figs. 4 and 5, fig. S8, and Table 1).

**SUPPLEMENTARY MATERIALS**

Supplementary material for this article is available at http://advances.sciencemag.org/cgi/content/full/5/11/eaav2869/DC1

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